

Review Packet

DIRECTIONS: Define the following terms.

Atom

Atomic number

Mass number

Subatomic particles

Proton

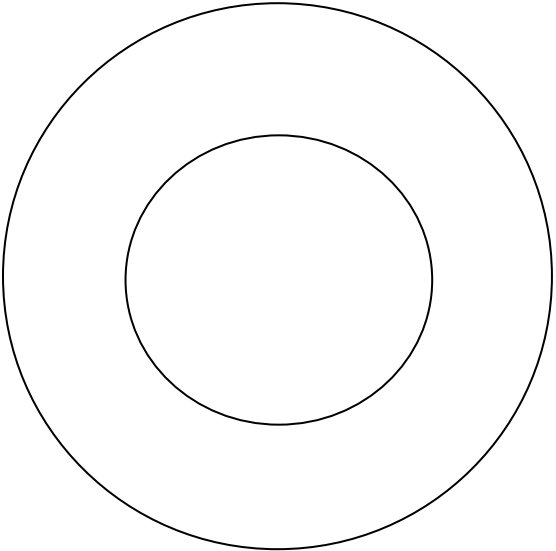
Electron

Neutron

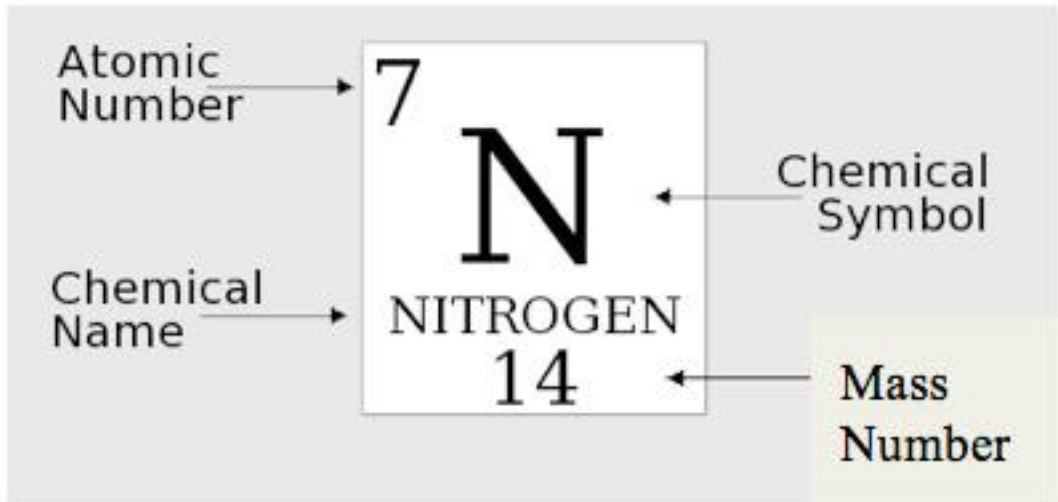
Nucleus

Electron cloud

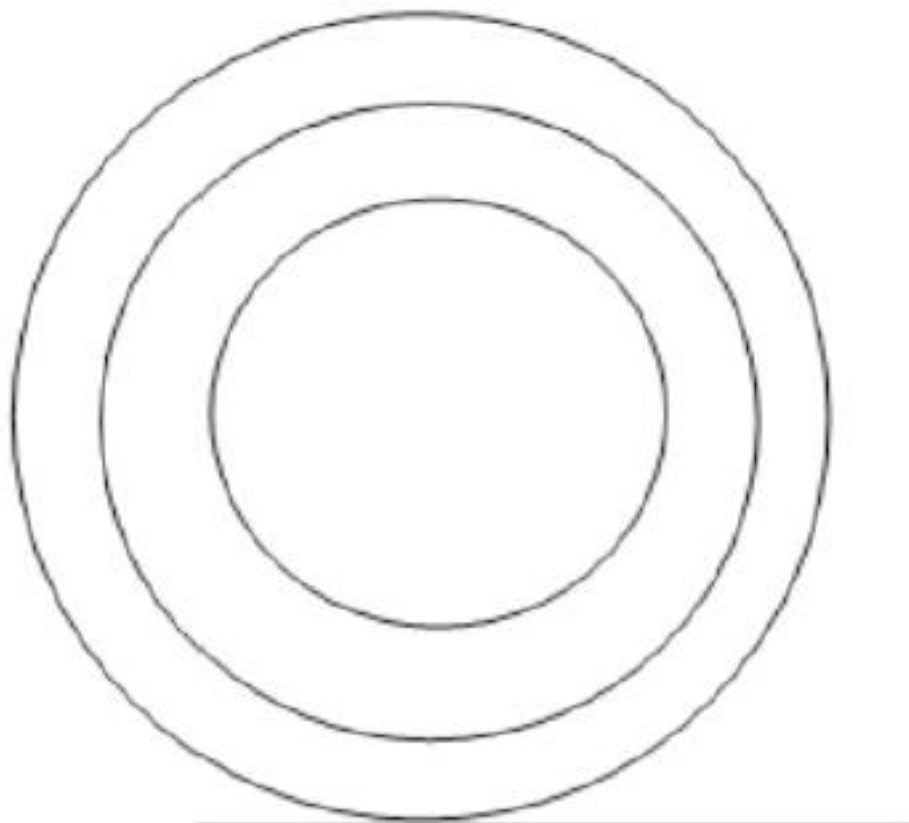
Energy levels

Define	Subatomic Particles		
<p>Nucleus: _____</p> <p>Atomic #: _____</p> <p>Mass #: _____</p> <p>Electron cloud: _____</p> <p>Maximum # of electrons (seats):</p> <ul style="list-style-type: none"> • Energy level 1 (inside): _____ • Energy level 2: _____ • Energy level 3: _____ 	Location	Charge	Name of Subatomic Particle
Show an Atom	Extension		
<p>Draw an atom below with:</p> <ul style="list-style-type: none"> • Atomic # = 6 • Mass # = 11 <div style="text-align: center; margin-top: 20px;">  </div>	<p>1) Create a rule to describe the relationship between positive and negative charges in an atom.</p> <p>_____</p> <p>_____</p> <p>_____</p> <p>_____</p> <p>_____</p> <p>2) All protons are identical. However, not all atoms are identical. Explain how this is possible.</p> <p>_____</p> <p>_____</p> <p>_____</p> <p>_____</p> <p>_____</p>		

This is a square from the periodic table of the elements:



Based on the information you see, draw an atom of Nitrogen below & fill in the chart:



Name _____

Date ____ Period ____

Section 3-3 & 3-4 HW

Directions: READ PAGES 68-71 – 3 times. Then, COMPLETE the questions below

1) pure substance - _____

- Give an example: _____

2) Element - _____

- Give an example: _____

3) Atom - _____

- Give an example: _____

4) Chemical symbol - _____

- Give an example: _____

Directions: READ PAGES 71-74 - 3 times. Then, COMPLETE the questions below.

1) Compound - _____

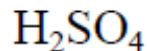
- Give an example: _____

2) Molecule - _____

- Give an example: _____

3) Chemical formula - _____

4) How many of each element are in this chemical formula:



H =

S =

O =